② 0.6 mol HF

$8 \cdot 10^{-2}$ mol NaOH $1L$
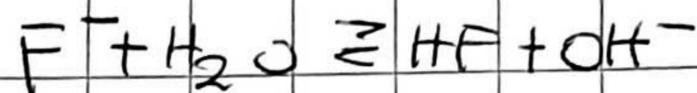
$\beta = ?$ $K_a = 3.3 \cdot 10^{-4}$

$NaOH \rightarrow Na^+ + OH^-$
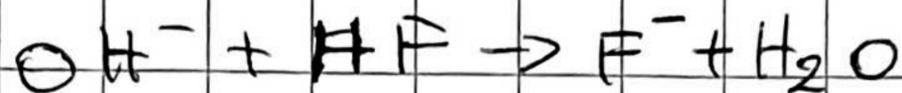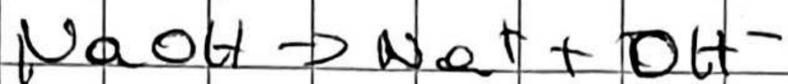
	OH^- +	$HF \rightarrow$	F^- +	H_2O
i)	$8 \cdot 10^{-2}$	0.6	/	—
f)	/	0.512	$8 \cdot 10^{-2}$	—

TAMPONE

	F^- +	$H_2O \rightleftharpoons$	HF +	OH^-
i)	$8 \cdot 10^{-2}$	—	0.512	/
eq)	$8 \cdot 10^{-2} - x$	—	$0.512 + x$	x

$K_b = \dfrac{[OH^-][HF]}{[F^-]}$

~~pppppppppeq~~

$\beta = 2.303 \left(\dfrac{C \, [H^+] K_a}{(K_a + [H^+])^2} \right)$

$2.86 \cdot 10^{-11} = \dfrac{x(0.512 + x)}{(8 \cdot 10^{-2} - x)}$

$[OH^-] = 4.46 \cdot 10^{-12}$

~~poooos pooo12 J~~

? $\boxed{pH = 2.65}$