② 0.6 mol HF

$8 \cdot 10^{-2}$ mol NaOH 1L

$\beta = ?$ $\quad K_a = 3.5 \cdot 10^{-4}$

$NaOH \rightarrow Na^+ + OH^-$

	OH^-	+	HF	$\rightarrow$	F^-	+	H_2O
i	$8 \cdot 10^{-2}$		0.6		/		-
f	/		0.512		$8 \cdot 10^{-2}$		-

$$\beta = 2.303 \left(\frac{C[H^+] K_a}{(K_a + [H^+])^2} \right)$$

TAMPONE

$F^- + H_2O \rightleftarrows HF + OH^-$

	F^-	+	H_2O	$\rightleftarrows$	HF	+	OH^-
i	$8 \cdot 10^{-2}$		-		0.512		/
eq	$8 \cdot 10^{-2} - x$		-		$0.512 + x$		x

$$K_b = \frac{[OH^-][HF]}{[F^-]}$$

$$2.86 \cdot 10^{-11} = \frac{x(0.512 + x)}{8 \cdot 10^{-2} - x} \approx$$

$[OH^-] = 4.46 \cdot 10^{-12}$

$pOH = 12.35$

$pH = 2.65$