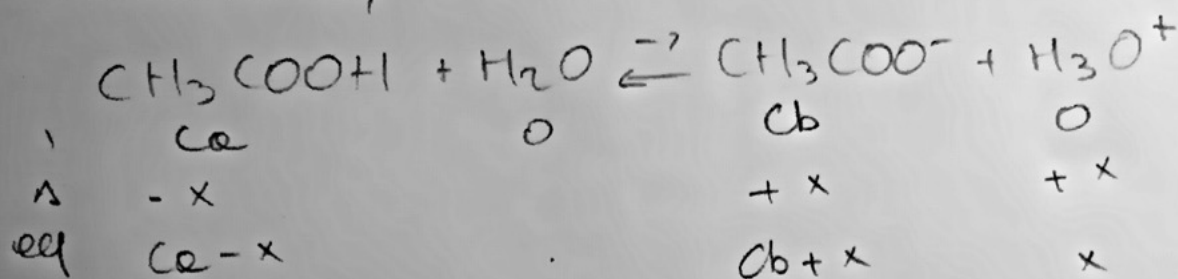
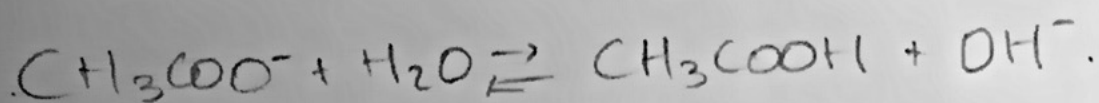
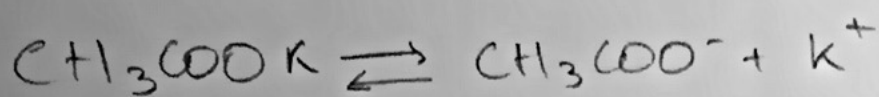
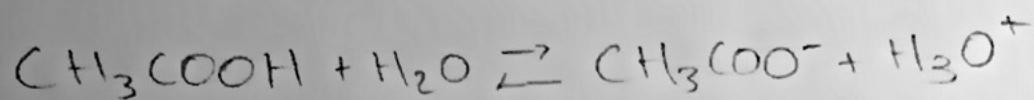


Ad un litro di soluzione di acido acetico
 $2,5 \times 10^{-2} [M]$ vengono aggiunti 1,7 g di acetato di
 Potassio CH_3COOK . Calcolare il pH (K_a acido acetico:
 $1,8 \times 10^{-5}$)

$$1,7g \times \frac{1 \text{ mol}}{98,15g} = 1,73 \times 10^{-2} \text{ moli} / 1L = 1,73 \times 10^{-2} [M] \text{ } \underline{CH_3COOK}$$



$$K_a = \frac{CH_3COO^- \times H_3O^+}{CH_3COOH} = \frac{(c_b + x) \times x}{c_a - x} ; \quad \begin{array}{l} c_b + x \approx c_b \\ c_a - x \approx c_a \end{array}$$

$$x = K_a \cdot \frac{c_a}{c_b} ;$$

$$pH = -\log x \Rightarrow -\log K_a - \left| \log \frac{c_a}{c_b} \right|$$

$$= (\log c_a - \log c_b)$$

$$pH = -\log K_a - \log c_a + \log c_b$$

$$pH = 0,58 ?$$