

④

Per ottenere una soluzione di:

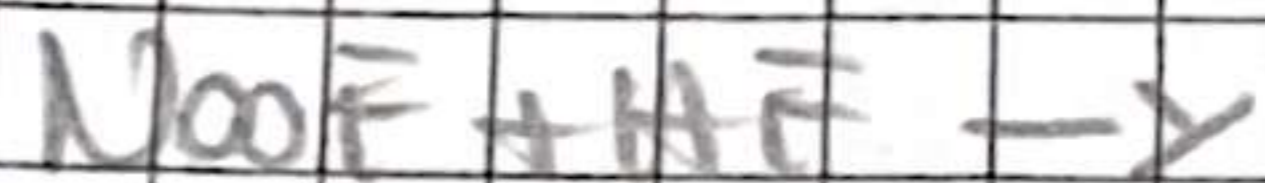
$$pH = 5,24$$

quantità di NaF (fluido de soda)

$$\text{sol: } V = 0,400 \text{ l}, 0,280 \text{ M de HA}$$

(K_a acido fluoridrico $2,10 \cdot 10^{-4}$)

SUCCINIMENTO



$$[H^+] = 10^{-pH} = 5,75 \cdot 10^{-6}$$

$$m_{HA} = 0,280 \cdot 0,400 \text{ l} = 0,112 \text{ mol}$$

$$m_{A^-} = K_a \cdot \frac{m_{HA}}{[H^+]}$$

$$m_{A^-} = \frac{126 \cdot 10^{-3} \cdot 0,112 \text{ mol}}{5,75 \cdot 10^{-6}} = 0,245 \text{ mol}$$

$$g = 0,245 \text{ mol} \cdot 41,98 \text{ g/mol} = 10,29 \text{ g}$$