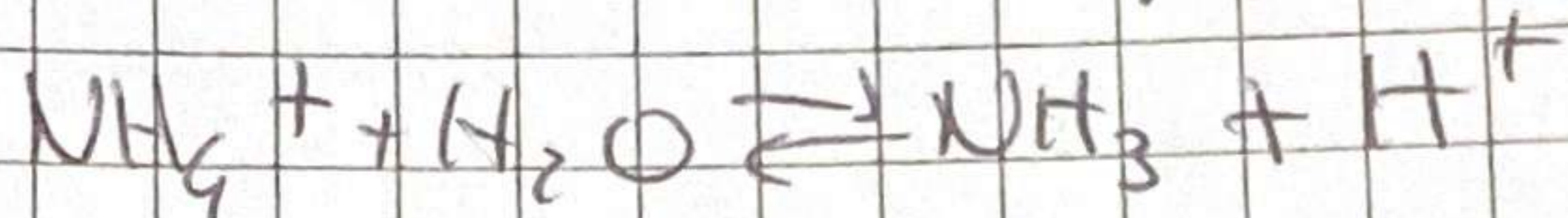
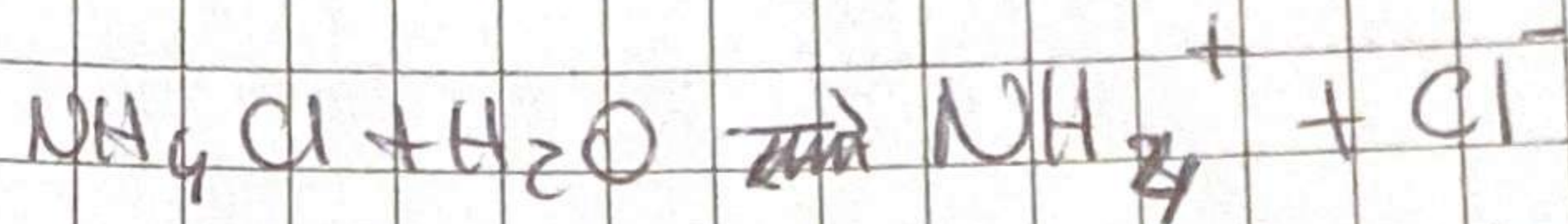


$\text{NH}_4\text{Cl}$  do idroksu azido



$$K_a = \frac{K_w}{K_b} = \frac{1,00 \cdot 10^{-14}}{1,8 \cdot 10^{-5}} = 5,55 \cdot 10^{-10} \cdot 0,432 = 2,39 \cdot 10^{-9}$$

$$[\text{H}^+] = \sqrt{2,39 \cdot 10^{-9}} = 4,88 \cdot 10^{-5}$$

$$\text{pH} = -\log(4,88 \cdot 10^{-5}) = 4,31$$