



$$n_{\text{HF}} = 0,100 \text{ mol} \cdot 0,120 \text{ mol} = 0,012 \text{ mol}$$

$$n_{\text{NaOH}} = 0,200 \text{ mol} \cdot 0,065 \text{ mol} = 0,013 \text{ mol}$$

$$\text{excess OH}^- = 0,013 - 0,012 = 0,001 \text{ mol}$$

$$[\text{OH}^-] = \frac{0,001 \text{ mol}}{0,300 \text{ L}} = \frac{0,001 \text{ mol}}{0,300 \text{ L}} = 3,33 \cdot 10^{-3}$$

$$\text{pOH} = -\log(3,33 \cdot 10^{-3}) = 2,47$$

$$\text{pH} = 14 - 2,47 = 11,52$$