$NaClO \rightleftarrows Na^+ + ClO^-$

$CaBr_2 \rightleftarrows Ca^{++} + 2Br^-$

$ClO^- + 2I^- \rightleftarrows Cl^- + I_2$

$I_2 + 2S_2O_4^- \rightleftarrows 2I^- + S_4O_6^=$

$IO_3^- + 5I^- \rightleftarrows 3I_2$

$AgNO_3 \rightleftarrows Ag^+ + NO_3^-$

$Ag^+ + Cl^- \rightleftarrows AgCl$

$Ag_{ecc} + SCN^- \rightleftarrows AgSCN$

$m\,AgNO_3 = 45\times10^{-3} \cdot 0{,}2\,M = 9{,}0\times10^{-3}$ mmol

$m\,NH_4SCN = 8\times10^{-3} \cdot 0{,}1 = 8\times10^{-4}$ mmol

$m\,Ag\,eff. = 9\times10^{-3} - 8{,}0\times10^{-4} = 8{,}2\times10^{-3}$ mmol $= m\,Cl^- = m\,ClO^-$ in 7 ml

$m\,NaClO$ in 70 ml $= 8{,}2\times10^{-3} : 7 = x : 70$ $\quad NaClO = 8{,}2\times10^{-2}$ mmol

$m\,NaClO = 8{,}2\times10^{-3} \cdot 74{,}4 = 6{,}10\,g$ $\quad \%\,NaClO = \frac{6{,}10}{10} \cdot 100 = 61\%$

$m\,KIO_3 = 48 : 100 = x : 15$ $\quad m\,KIO_3 = 7{,}2\,g$ $\quad m\,KIO_3 = \frac{7{,}2}{214} = 0{,}033$ mmol

$m\,I_2 = 0{,}033 : 1 = x : 3$ $\quad m\,I_2 = 0{,}1$ mmol

$m\,Na_2S_2O_3 = 0{,}1 : 1 = x : 2$ $\quad x = 0{,}2$ mmol $= m\,Na_2S_2O_3$ in 24,1 ml

$m\,Na_2S_2O_3$ in 17 ml $= 0{,}2 : 24{,}1 = x : 17$ $\quad m\,Na_2S_2O_3 = 0{,}14$ mmol

$m\,I_2 = 0{,}14 : 2 = x : 1$ $\quad m\,I_2 = 0{,}07$ mmol