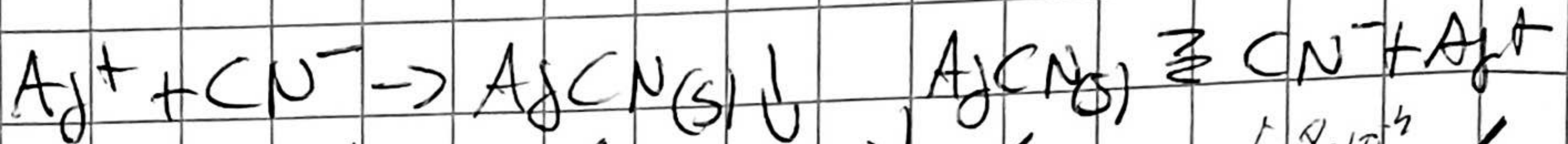
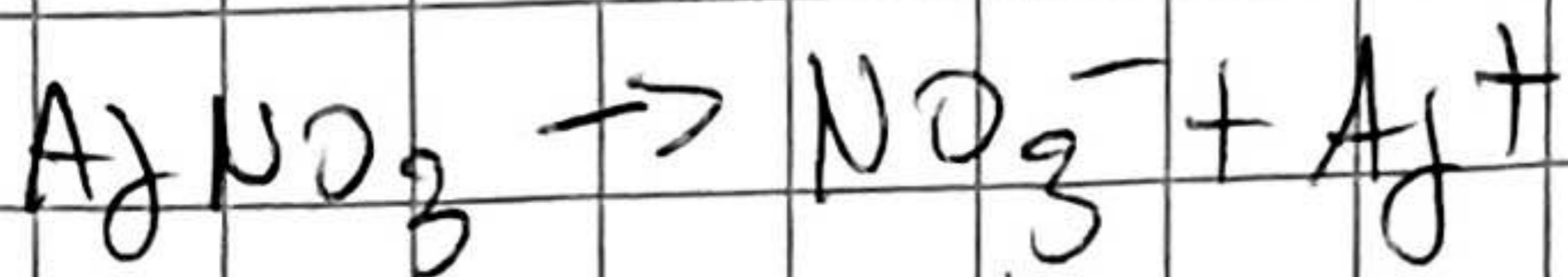


AgCN ? 22 mL 10^{-2} M AgNO_3 $C_f = 4.7 \cdot 10^{-4}$
 +
 450 mL 10^{-3} M NaCN $C_f = 9.5 \cdot 10^{-4}$

$\rho_{\text{AgCN}} = 133.89 \text{ g/mol}$

$K_{sp} \text{AgCN} = 5.57 \cdot 10^{-17}$



\downarrow \downarrow
 $4.7 \cdot 10^{-4}$ $9.5 \cdot 10^{-4}$
 \downarrow \downarrow
 $4.9 \cdot 10^{-4}$ $4.7 \cdot 10^{-4}$

\downarrow \downarrow
 $4.7 \cdot 10^{-4} - s$ $4.8 \cdot 10^{-4} + s$

$K_{sp} = (4.8 \cdot 10^{-4} + s) \cdot s$

$s^2 + 4.8 \cdot 10^{-4} s - K_{sp} = 0$

$s = 1.24 \cdot 10^{-13}$

$\rho_{\text{AgCN}} = 4.7 \cdot 10^{-4} - x = 4.7 \cdot 10^{-4}$

$\rho_{\text{AgCN}} = 4.7 \cdot 10^{-4} \cdot 133.89 \text{ g/mol}$

$= 6.28 \cdot 10^{-2} \text{ g/L}$

$\rho = 6.28 \cdot 10^{-2} \text{ g/L} \cdot 0.4744$

$\rho = 2.98 \cdot 10^{-2} \text{ g/L}$