

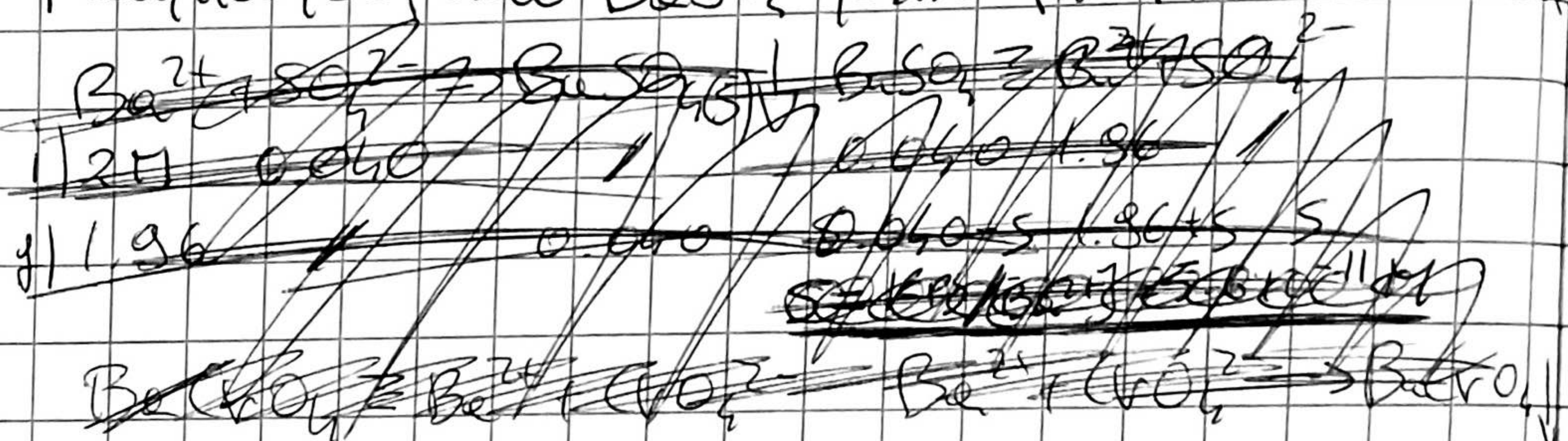
a) quale precipita per primo

K_{ps} più piccolo

$K_{ps} \text{BaSO}_4 = 1.1 \cdot 10^{-10}$

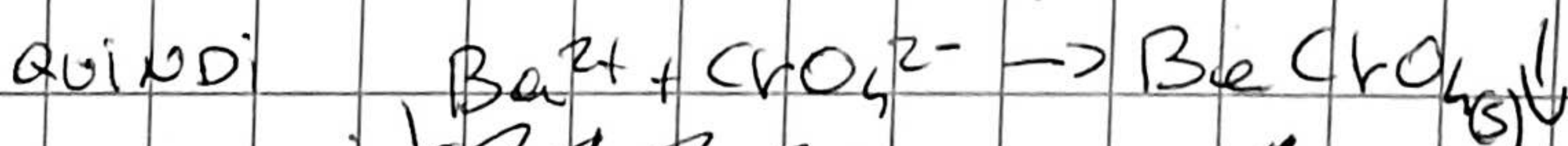
$K_{ps} \text{BaCrO}_4 = 2.1 \cdot 10^{-10}$

precipita per primo BaSO_4 perché più piccolo di BaCrO_4



• INCOGNITA $\text{BaCrO}_4 \rightleftharpoons \text{Ba}^{2+} + \text{CrO}_4^{2-}$

CONSIDERO 99% precipitato BaSO_4 e quindi 0.01% di $\text{Ba}^{2+} \rightarrow 2 \cdot 10^{-4} \text{ M}$



~~$1.2 \cdot 10^{-4}$~~ 0.020
~~1.98 \cdot 10^{-3}~~ $2 \cdot 10^{-4}$

$K_{ps} = [\text{CrO}_4^{2-}][\text{Ba}^{2+}] \rightarrow [\text{CrO}_4^{2-}] = \frac{K_{ps}}{[\text{Ba}^{2+}]} = 1.05 \cdot 10^{-6}$