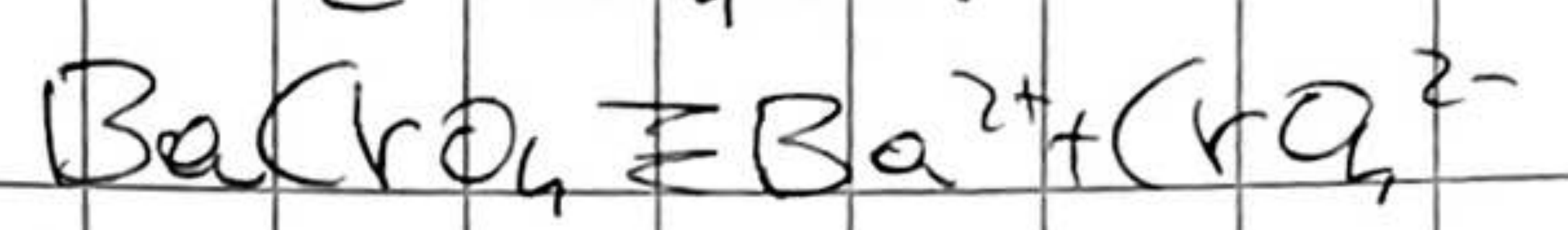
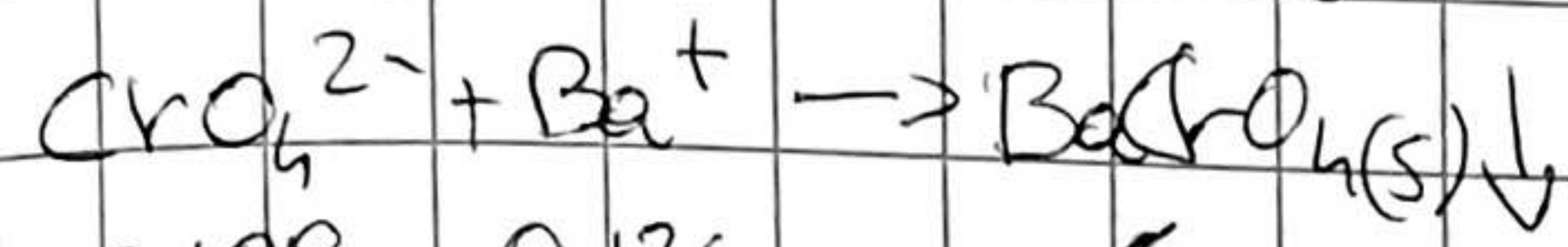
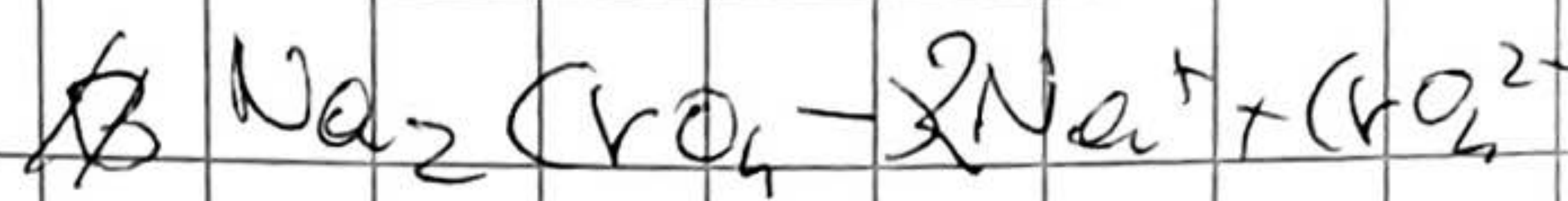
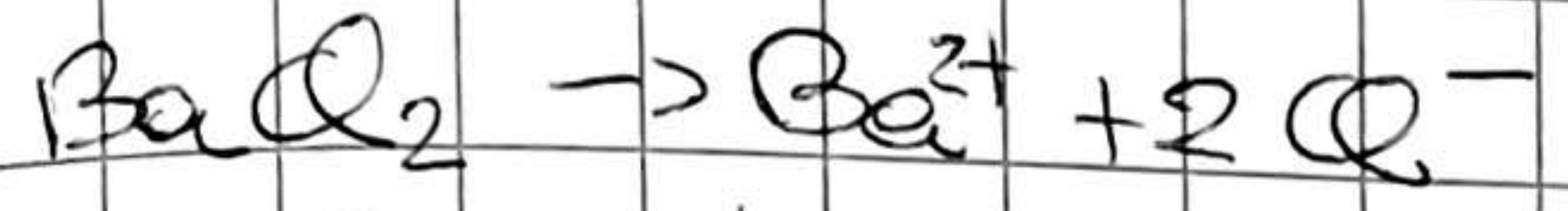


3) 25 mL 0.300M BaCl<sub>2</sub> + 30 mL Na<sub>2</sub>CrO<sub>4</sub> 0.2M  
 BaCrO<sub>4</sub>; g BaCrO<sub>4</sub> precipitato e concentrazioni residue Ba<sup>2+</sup>; CrO<sub>4</sub><sup>2-</sup>

$$C_1 V_1 = C_2 V_2 \rightarrow C_2 = C_1 \frac{V_1}{V_2} = 0.136 M$$

BaCl<sub>2</sub>

$$Na_2CrO_4 - C_2 = 0.108 M$$



i)	0.108	0.136	/
ii)	/	0.027	0.108

i)	0.108	0.027	/
ii)	0.09 - s	0.027 + s	s

$$K_{PS} = [Ba^{2+}][CrO_4^{2-}]$$

$$K_{PS} = s^2 \rightarrow s = \sqrt{K_{PS}} = 1.55 \cdot 10^{-5}$$

$$g BaCrO_4 = 0.108 - 1.55 \cdot 10^{-5} = 0.108$$

$$g BaCrO_4 = 0.108 \text{ mol} \cdot 253.33 \frac{g}{\text{mol}}$$

$$= 27.6 \frac{g}{\text{mol}} \cdot 0.055 \frac{\text{mol}}{L}$$

$$g = 0.835 g$$

La quantità residue somma

$$[Ba^{2+}] = K_{PS} = 0.027 M$$

$$[CrO_4^{2-}] = 1.55 \cdot 10^{-5} M$$