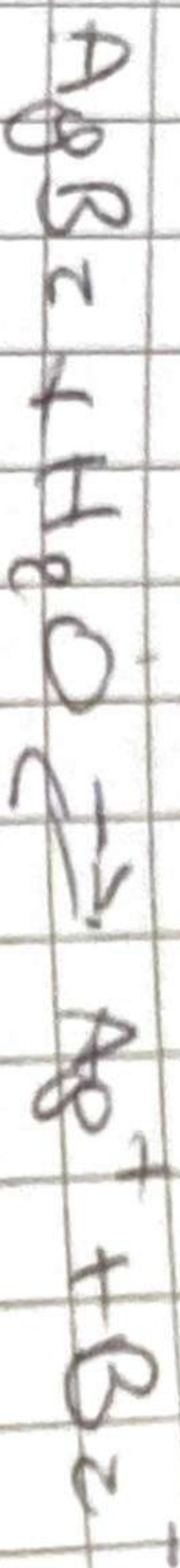


SOLUCIÓE



$$K_{ps} = [Ag^+]^2 [CrO_4^{2-}]$$

$$K_{ps} = S^3$$

$$5,0 \cdot 10^{-13} = S^3$$

$$S = \sqrt[3]{5,0 \cdot 10^{-13}} = 7,937 \cdot 10^{-5}$$

$$p = 7,937 \cdot 10^{-5} \quad \text{ITZ ITZ phor} = 4,32 \cdot 10^{-4} \text{ g}$$

$$\text{um } 350 \text{ ml} = \frac{4,32 \cdot 10^{-4} \cdot 0,35 \text{ l}}{1,00 \text{ l}} = 1,512 \cdot 10^{-4} \text{ g}$$

in  $CrO_4^{2-}$  ho

$$K_{ps} = [Ag^+]^2 [CrO_4^{2-}]$$

$$K_{ps} = [CrO_4^{2-}] [Ag^+]^2 = 0,5520 [Ag^+]^2$$

$$5,0 \cdot 10^{-13} = S \cdot (5 + 0,5520)$$

$$S = \frac{5,0 \cdot 10^{-13}}{5,5520} = 8,99 \cdot 10^{-14}$$

$$8 = 277 - 12 \cdot \Delta 87 \cdot \Delta T_{mod} = 25 \cdot 20 - 12 \cdot 20$$

$$K_m \cdot 350 \text{ mm} = 5 \cdot 20 - 12 \cdot 20 \cdot 0.352 = 1,82 \cdot 20$$

1.2